

Experiment: Taking Crucibles to Constant Weight

Name: _____

Date: _____

SS #: _____

Balance # _____

Date	Crucible 1	Difference	Crucible 2	Difference	Crucible 3	Difference
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_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____
_____	_____	_____	_____	_____	_____	_____

Constant Wt.

Crucible 1

Crucible 2

Crucible 3

Experiment: Review of Laboratory Techniques

Name: _____

Date: _____

SS #: _____

1. Calibration of Your 25 mL Pipet:

Water Temperature: _____

	1	2	3
Wt of Volumetric Flask + Water:	_____	_____	_____
Wt of Volumetric Flask:	_____	_____	_____
Wt of Water:	_____	_____	_____
	4	5	6
Wt of Volumetric Flask + Water:	_____	_____	_____
Wt of Volumetric Flask:	_____	_____	_____
Wt of Water:	_____	_____	_____

2. Preparation of a Standard KCl Solution

Wt of Wt. Bottle + KCl(Original):	_____	_____	_____
Wt of Wt. Bottle + KCl(Remain):	_____	_____	_____
Wt of KCl removed:	_____	_____	_____

3. Use of Buret and Preparation of a Diluted KCl Solution

Final Buret Reading:	_____	_____
Initial Buret Reading:	_____	_____
Volume Delivered:	_____	_____

Experiment: Gravimetric Determination of Nickel by Homogeneous Precipitation with Dimethylglyoxime

Name: _____

Date: _____

SS: _____

Unknown #: _____

Balance # _____

Urea:

Beaker: _____

Beaker + Urea: _____

Total Urea: _____

Date	Crucible 1	Crucible 2	Crucible 3
	Difference	Difference	Difference
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____
_____	_____	_____	_____

Constant Wt. (Crucible + Precipitate)

Crucible 1	Crucible 2	Crucible 3
_____	_____	_____

Constant Wt. (Crucible, from Taking Crucibles to Constant Weight)

Crucible 1	Crucible 2	Crucible 3
_____	_____	_____

Experiment: Acid-Base Titration-HCl standardization

Name: _____

Date: _____

SS #: _____

Unknown #: _____

1. Standardization of the 0.1M HCl Solution:

	1	2	3
Wt of Wt. Bottle + Na ₂ CO ₃ (Original):	_____	_____	_____
Wt of Wt. Bottle + Na ₂ CO ₃ (Remain):	_____	_____	_____
Wt of Na ₂ CO ₃ removed:	_____	_____	_____
	4	5	6
Wt of Wt. Bottle + Na ₂ CO ₃ (Original):	_____	_____	_____
Wt of Wt. Bottle + Na ₂ CO ₃ (Remain):	_____	_____	_____
Wt of Na ₂ CO ₃ removed:	_____	_____	_____

Titration

	1	2	3
Final Buret Reading:	_____	_____	_____
Initial Buret Reading:	_____	_____	_____
Volume Delivered:	_____	_____	_____
	4	5	6
Final Buret Reading:	_____	_____	_____
Initial Buret Reading:	_____	_____	_____
Volume Delivered:	_____	_____	_____

Experiment: Acid-Base Titration-Unknown

2. Determination of the Unknown Carbonate:

	1	2	3
Wt of Wt. Bottle + Unknown (Original):	_____	_____	_____
Wt of Wt. Bottle + Unknown (Remain):	_____	_____	_____
Wt of Unknown removed:	_____	_____	_____
	4	5	6
Wt of Wt. Bottle + Unknown (Original):	_____	_____	_____
Wt of Wt. Bottle + Unknown (Remain):	_____	_____	_____
Wt of Unknown removed:	_____	_____	_____

Titration

	1	2	3
Final Buret Reading:	_____	_____	_____
Initial Buret Reading:	_____	_____	_____
Volume Delivered:	_____	_____	_____
	4	5	6
Final Buret Reading:	_____	_____	_____
Initial Buret Reading:	_____	_____	_____
Volume Delivered:	_____	_____	_____

**Experiment: The Use of Fluoride Ion Selective Electrode for the
Determination of Fluoride in Drinking Water and Unknown Solution**

Concentration of Standard Solution				
mV				

Unknown #:

	Unknown 1	Unknown 2	DD water	Drinking water
mV				

Experiment: The Spectronic 20 Spectrophotometer & Spectrophotometric Determination of Manganese in Steel

Conc. Of Original KMnO_4 :

Volume of original solution used:

Total diluted volume:

Concentration of dilute KMnO_4 :

Delivered mL					dilute KMnO_4
Abs					

Unknown #:

Volume of pipet:

	Unknown 1	Unknown 2
Absorption Reading:	_____	_____
	_____	_____

Experiment: Antimony Determination

Name: _____

Date: _____

SS #: _____

Unknown #: _____

Standardization of I_2 :

mL of I_2 pipetted: _____

Normality of Na_2SO_3 : _____ N

1

2

3

mL of Na_2SO_3 titrated _____

Normality of I_2 : _____

Average of Normality of I_2 : _____

Average Deviation: _____

Relative Average Deviation: _____

Determination of Sb:

1

2

3

mL of Sb pipetted: _____

mL of I_2 titrated _____

Concentration of Sb in 100 mL volumetric flask (N) _____

Concentration of Sb in 100 mL volumetric flask (M) _____

Concentration of Sb in 100 mL volumetric flask (ppm) _____

Average of Concentration of Sb in 100 mL volumetric flask (ppm): _____

Average Standard Deviation: _____

CV%: _____